Name: ­­\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Per: \_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_

Classifying Chemical Reactions

More Practice!

**Part I:**

Classify each of the following reactions as Synthesis (S), Decomposition (D), Single Replacement (SR), Double Replacement (DR), or Combustion (C).

1. \_\_\_\_\_\_\_\_ BaCl2 + Na2 SO4 → 2 NaCl + BaSO4
2. \_\_\_\_\_\_\_\_ C3H8 + 5O2 → 3 CO2 + 4 H2O
3. \_\_\_\_\_\_\_\_ 3 NH3 + 3I2 → 2 NI3 + 3 H2
4. \_\_\_\_\_\_\_\_ 2 H2O2 → 2 H2O + O2
5. \_\_\_\_\_\_\_\_ Cl2 + 2 NaBr → 2 NaCl + Br2

**Part II:**

Predict the products of the following reactions. Remember to make sure you balance each equation. Reminders: oxidation states, diatomic vs. monatomic molecules, and solubility rules.

1. S8 + F2 →
2. HNO3 + KOH →
3. Al2O3 →
4. CuO +H2 →
5. PbSO4 + AgNO3 →